

# Chapter 1: "Opener and Review" Worksheet

1. A solution with a pH between 0 and 6 would be: Circle One.

Acidic

Basic

2. A solution with a pH between 8 and 12 would be: Circle One.

Acidic

Basic

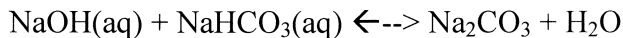
3. Indicate the direction of the dipole moment, if any. Indicate whether the compound is polar or non-polar.



4. Draw a Lewis dot structure for each of the following:



5. Rewrite the following equations to show the Lowry-Bronsted acids and bases actually involved. Label each as stronger or weaker.



6. A neutral (formal charge 0) atom of:

C always has 4 bonds to it.



H always has 1 bonds to it.



N always has 3 bonds to it.



O always has 2 bonds to it.

